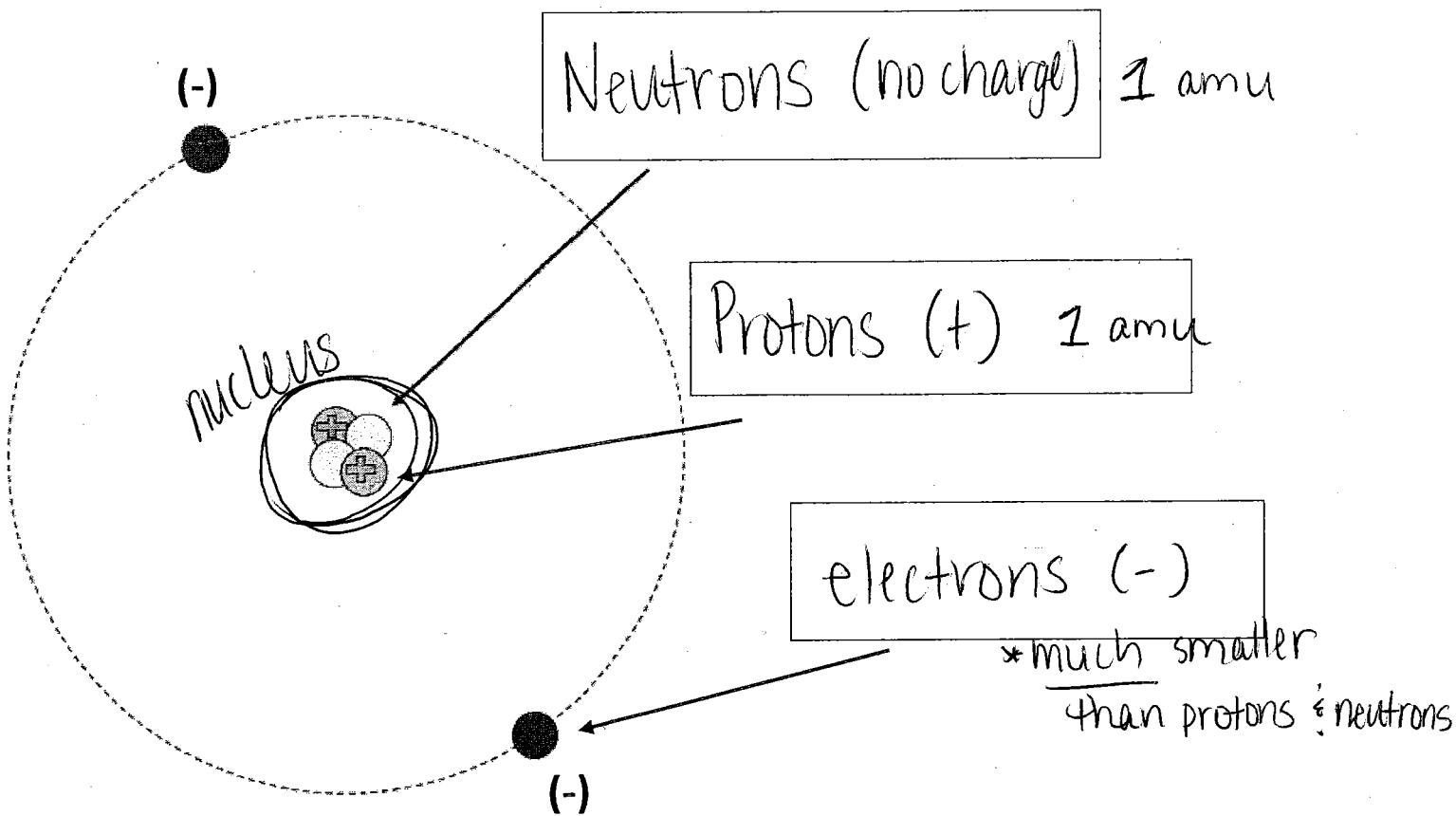


Atom Notes

An **ATOM** is the building block of matter



| | | |
|---------|---|--|
| 14 | ← | <u>Atomic number</u> (= # of protons) |
| Si | ← | <u>Chemical symbol</u> |
| 28.086 | ← | <u>Atomic mass</u> |
| Silicon | ← | <u>element name</u> |

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APE



A atomic number

= (equals)

P rotons

= (equals)

E lectrons

MAN



M ass

- (minus)

A atomic number

= (equals)

N eutrons

**ATOMIC
MASS**

—

**ATOMIC
NUMBER**

=

NEUTRONS

Atomic #

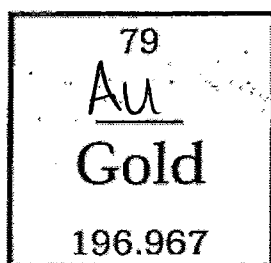
=

Protons

=

Electrons

Let's Practice!



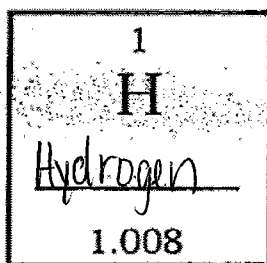
Atomic # = 79

Atomic Mass = 197

of Protons = 79

of Neutrons = 118

of Electrons = 79



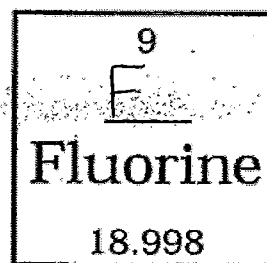
Atomic # = 1

Atomic Mass = 1

of Protons = 1

of Neutrons = 0

of Electrons = 1



Atomic # = 9

Atomic Mass = 19

of Protons = 9

of Neutrons = 10

of Electrons = 9